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Depression in freezing point Concept in 12 minutes QUICK SUMMARY Colligative property

Introduction to Explore Learning Gizmos Solutions (Part 6) - Colligative Properties | Class 12 - NCERT Class 12th Chemistry : Numericals on Colligative Property- Part 1 Colligative Properties - Boiling Point Elevation and Freezing Point Depression Solutions 08 | Van't Hoff Factor and Abnormal Molar Masses - Most Important Concept IIT JEE/NEET Solutions 03 | Relative Lowering of Vapour Pressure due to Non Volatile Solute Colligative Property

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Determine how the physical properties of a solvent are dependent on the number of solute particles present. Measure the vapor pressure, boiling point, freezing point, and osmotic pressure of pure water and a variety of solutions. Compare the effects of four solutes (sucrose, sodium chloride, calcium chloride, and

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potassium chloride) on these physical properties.

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Colligative Properties. Launch Gizmo. Determine how the physical properties of a solvent are dependent on the number of solute particles present. Measure the vapor pressure, boiling point, freezing point, and osmotic pressure of pure water and a variety of solutions. Compare the effects of four solutes (sucrose, sodium chloride, calcium chloride, and potassium chloride) on these physical properties.

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Colligative property. a property of a solution that depends on the number of particles of solute in a given volume of solvent. o Colligative properties include boiling point, freezing point, vapor

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____ Gizmo Warm-up Adding salt or other substances to water can affect the temperature at which it freezes or boils. These effects and others, known as colligative properties, are explored in the Colligative Properties Gizmo. Check that No solute is selected and the Air temp. is 25 ° C. 1.

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Gizmo Warm-up Adding salt or other substances to water can affect the temperature at which it freezes or boils. These effects and others, known as colligative properties, are explored in the Colligative Properties Gizmo™. Check that No solute is selected and the Air temp is 25 ° C. 1. Describe the motion of the water molecules in the beaker.

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This question and others are explored in the Colligative Properties Gizmo. In this Gizmo, students can add different substances to water and see how they affect its freezing point, boiling point, and osmotic pressure. The key is realizing that the presence of solute particles decreases the number of collisions between water molecules, thus decreasing the chances of bonds forming between them.

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a property of a solution that depends on the number of particles of solute in a given volume of solvent. o Colligative properties include boiling point, freezing point, vapor pressure, and osmotic pressure. o Colligative properties depend on the concentration of solute as well as the number of particles a molecule of solute dissociates into.

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