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successful. As

understood, capability

does not recommend

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capably as covenant

even more than

additional will

provide each success.

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sending people to

hell? /" Pearson

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reading the section in
your textbook,
respond to each
statement. 1. State
the term for the
amount of energy
that is needed for a

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chemical reaction to occur. 2. Summarize the relationship between an enzyme and a substrate. enc 3. <https://www.svcasd.org/cms/lib07/NY01913388/Centricity/Domain/245/biochem%20review.pdf>.

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must conserve the
balance of mass.

identify which

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number indicates the number of atoms of each element in a molecule. the number 6. minimum amount of energy required for reactants to form products. activation energy. substance that lowers energy needed to start a chemical reaction. catalyst.

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but ionic-crystalline compounds are not. The metallic bond is the same in all directions throughout the metallic structure allowing the atoms to slide past each other. This sliding is why metals are ductile and malleable.

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chemical reaction.

reactant. product.

Bonding
activation energy. the

process by which

atoms or groups of

atoms in substances

ar.... substance that

exists before a

chemical reaction

starts; locat....

substance formed by

a chemical reaction;

located on the

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Class In your

textbook, read about
reactants and

products. Fill in the
blanks with the

correct number of
molecules to balance
the chemical

equation. $C_6H_6 + 12$

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6 (1) Respond to each statement. $\text{CO}_2 + \text{H}_2\text{O}$
(3) (2) 4. State the principle that explains why there must be the same number of atoms of

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Chemical Bonding

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ANSWER Answer the following questions in the space provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond approach each other, the

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6 Section 2 Chemical Bonding

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positive ion.

Molecules. a neutral group of atoms that are held together by covalent bonds.

Molecular compound. a chemical compound whose simplest units are molecules.

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Chemical formula.
indicates the relative
... Chemical Bonding:
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Covalent Answer Key
5. 2. 6. 3. Section 1
and 3 (page 114) 1.
Second Law of
Motion. The force
necessary to move

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the space shuttle is equal to its mass times its acceleration.

2. First Law of Motion. The rock stays at rest at the top of the hill until the boy applies force that sends it rolling downhill. Sections 2 (page 115) 1. gravity 6. attraction. 2 ...

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PLAY. Match. Gravity.
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set (10) Chemical
reaction. Process in
which atoms or

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groups of atoms in
substances are
reorganized into
different substances.

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Chemical Reactions ...

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Composition 1. 100

washers 0.110 g 1

washer = 11.0 g

(assuming 100

washers is exact) 100.

g 1 washer 0.110 g =

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909 washers 2. 500.
g 1 cork 1.63 g =
306.7 = 307 corks
500. g 1 stopper 4.31
g = 116 stoppers 1 kg
(1000 g) of corks
contains (1000 g 1
cork

Chapter 6 Chemical
Composition - Francis
Howell High School
c. In a chemical
equation, the

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Section 2
Chemical
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reactants are on the right, and the products are on the left. d. When balancing a chemical equation, you can never change the subscripts of any chemical formula. e. In chemical reactions, matter is neither created nor destroyed, so a chemical equation

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must have the same number of atoms on both sides of the equation.

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editorial corrections

Table 6.2-2 was

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updated... 2

Chemical

Chapter 6: Organic

Chemical Process

Industry, AP 42, Fifth

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6.2 Chemical

Sedimentary Rocks.

Whereas clastic

sedimentary rocks are

dominated by

components that have

been transported as

solid clasts (clay, silt,

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sand, etc.), chemical sedimentary rocks are dominated by components that have been transported as ions in solution (Na^+ , Ca^{2+} , HCO_3^- , etc.). There is some overlap between the two because almost all clastic sedimentary rocks contain cement formed from dissolved ions, and

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many chemical
sedimentary rocks
include some clasts.

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