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Chapter 4 Aqueous Reactions and Solution Stoichiometry. Solutions: •Homogeneous mixtures of two or more pure substances. •The solvent is usually present in greatest abundance. •Or, the solvent is the liquid when a solid is dissolved •All other

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substances are solutes. Dissociation.

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- Homogeneous mixtures of two or more pure substances.
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Aqueous Reactions.

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~~Chapter 4 Aqueous Reactions and Solution Stoichiometry~~

Chapter 4. Reactions in Aqueous Solution. James F. Kirby. Quinnipiac University. Hamden, CT. Solutions. Solutions are defined as homogeneous mixtures of two or more pure substances. The . solvent. is present in greatest abundance. All other substances are . solutes. When water is the solvent, the

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solution is called an .

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Section 4.4 - Oxidation-Reduction Reactions.
Oxidation-reduction reactions (redox reactions) are reactions in which electrons are transferred between reactants. Oxidation is loss of electrons. Reduction is gain of electrons. Oxidation and reduction always occur together.

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Chapter 4. Aqueous Reactions and Solution Stoichiometry
4.1 General Properties of Aqueous Solutions

- A solution is a homogeneous mixture of two or more substances.
- A solution is made when one substance (the solute) is dissolved in another (the solvent).
- The solute is the substance that is present in the smallest amount.

~~Brown LeMay Chapter 4.pdf — Chapter 4 Aqueous Reactions ...~~

Chemistry Chapter 4: Aqueous Reactions and Solution Stoichiometry. Elemental form of

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atom is 0 because it has not gained or lost anything; monatomic ions are equal to their charge; nonmetals in compounds are negative [F is always -1 and O is always -2; H can be positive or negative]; sum of all numbers in a neutral compound is 0 and in a polyatomic ion it is the overall charge of the ion.

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The acidity or basicity of an aqueous solution is described quantitatively using the pH scale. 4.9: Oxidation-Reduction Reactions Oxidation-reduction reactions are balanced by separating the overall chemical equation into an oxidation equation and a reduction equation.

~~4: Chemical Reactions and Aqueous Reactions Chemistry ...~~

Chapter 4. Aqueous Reactions and Solution

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Stoichiometry Common Student Misconceptions •
Molarity is moles of solute per liter of solution, not per liter of solvent. •
Students sometimes use moles instead of molarity in $M_{\text{initial}} V_{\text{initial}} = M_{\text{final}} V_{\text{final}}$. •
Students sometimes think that water is a good conductor.

~~Chapter 4. Aqueous Reactions and Solution Stoichiometry~~

Larson-Foothill College 1 Chapter 4: Aqueous Reactions and Solution Stoichiometry The topics in this chapter will further our knowledge of types of chemical reactions and

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our abilities to predict the products of and write balanced chemical equations for a variety of chemical reactions.

~~Chapter 4a Chapter 4 Aqueous Reactions and Solution ...~~

Home assignment: Chapter 4: Aqueous Reactions and Solution Stoichiometry 7) In which species does sulfur have the highest oxidation number? A) S, (elemental form of sulfur) B) H₂S C) SO₂ D) H₂SO₄ E) K₂SO₄, 8) One method for removal of metal ions from a solution is to convert the metal to its elemental form so it can be filtered out as a

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solid.

~~Solved: Home Assignment: Chapter 4: Aqueous Reactions And ...~~

Chemistry, The Central Science, 10th edition Theodore L. Brown; H. Eugene LeMay, Jr.; and Bruce E. Bursten Chapter 4 Aqueous Reactions and Solution Stoichiometry - PowerPoint PPT presentation. Water is the dissolving medium, or solvent. Water is bent or V-shaped. The O-H bonds are covalent.

~~PPT - Chapter 4 Aqueous Reactions and Solution ...~~

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Chapter 4 Worksheet Spring 2007 page 4 of 4
Complete, balance, and identify the reaction type for each of the following equations:

Type 18. $\text{MgO (s)} + \text{H}_2\text{O (l)} \rightarrow \text{Mg(OH)}_2 \text{ (s)}$
combo. 19. $\text{Zn (s)} + \text{Cu(NO}_3)_2 \text{ (aq)} \rightarrow \text{Zn(NO}_3)_2 \text{ (aq)} + \text{Cu (s)}$ SR, metal
20. $\text{Ba(NO}_3)_2 \text{ (aq)} +$

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$\text{MgSO}_4 (\text{aq}) + \text{BaSO}_4 (\text{s}) + \text{Mg}(\text{NO}_3)_2 (\text{aq})$
precipitation 21.

~~Chapter 4 Practice Worksheet: Reactions in Aqueous Solutions~~

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