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Molecule: a Review. •

Water is a simple tri-
atomic molecule, H. 2.

O. •Each O-H bond is
highly polar, because

of the high

electronegativity of

the oxygen (N, O, F,

and Cl have high

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values) • bond angle
of water = 105° .

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Page 501: 11 Answer
All ionic compounds

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are electrolytes because they dissociate into ions and conduct an electrical current in aqueous solution or in the molten state.

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chaotic movement of colloidal particles, caused by collision with particles of the solvent in which they are dispersed. a compound that conducts an electric current when it is in an aqueous solution or in the molten state; all ionic compounds are electrolytes, but most covalent

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Compounds are not.

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heterogeneous

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mixture containing particles that are small enough to remain dispersed in the solvent and do not separate on standing.

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surface tension. **Test B**

surfactant. aqueous
solution. solvent. the
inward force or pull
that tends to minimize
the surface ar.... any
substance that
interferes with
hydrogen bonding
between wa.... is water

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that contains
dissolved substances.
the dissolving medium
in a solution.

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WATER AND ITS

PROPERTIES (pages
445–449) This section

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Chapter 15 describes the properties of water in the liquid and solid states and explains how hydrogen bonding affects the surface tension and vapor pressure of water. Water in the Liquid State (pages 445–447) 1.

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PROPERTIES

(pages 445–449)

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answers , source:eur-
lex.europa.eu He may
want to stretch

himself once a worker
knows his efforts do
not go unnoticed. For
instance, if he knows
his performance will

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be judged based on achievement of a target, he will work harder to achieve it.

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crystal of solid NaCl is placed into an aqueous NaCl solution. No

precipitate forms in the bottom of the container. The final solution is: a.) super saturated ... Chapter 5: Soil Water 57 Terms.

hannah_lange5.

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– Explain the high surface tension and low vapor pressure of water in terms of the structure of the water molecule and hydrogen bonding.

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separate on standing.

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Water World 15.1.
Collect and Organize
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four lines to describe
the possible
dependence of

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percent ionization of acetic acid with concentration. We are to choose the one that best represents the trend for this weak acid. Analyze

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